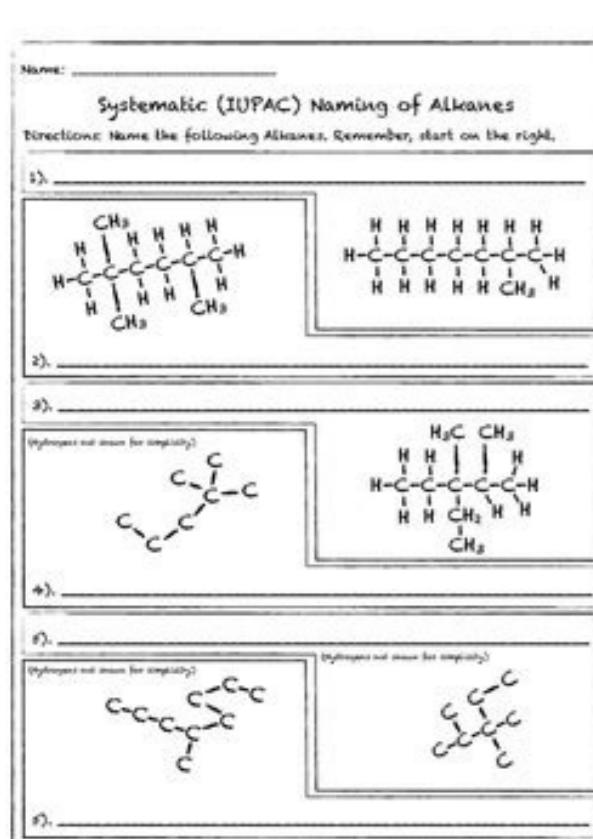


I'm not a robot!

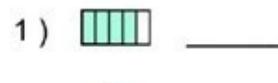
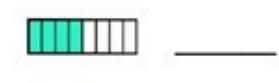
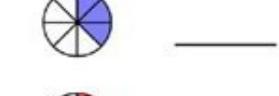
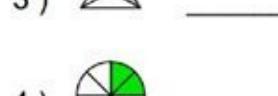
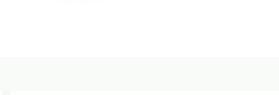




Solutions for the Naming Ionic Compounds		(Still) More Naming Practice - Answers	
<b>Practice Worksheet</b>		<b>Answers</b>	
1) ammonium chloride	22) iron (II) phosphate	1) BBr <sub>3</sub>	boron trisulfide
2) iron (II) nitrate	23) titanium (II) selenide	2) CaSO <sub>4</sub>	calcium sulfate
3) titanium (III) bromide	24) calcium iodide	3) C <sub>2</sub> Br <sub>6</sub>	dicarbon hexabromide
4) copper (I) phosphide	25) sodium chloride	4) Cr(O <sub>3</sub> ) <sub>3</sub>	chromium (VI) carbonate
5) tin (IV) sulfide	26) zinc carbonate	5) Ag <sub>3</sub> P	silver phosphate
6) gallium arsenide	27) iron (II) amosite	6) VO <sub>2</sub>	iodine dioxide
7) lead (IV) sulfate	28) copper (II) chromite	7) PbS	lead (II) sulfide
8) beryllium bicarbonate	29) ammonium oxide	8) CH <sub>4</sub>	methane
9) manganese (II) sulfide	30) potassium hydroxide	9) N <sub>2</sub> O <sub>3</sub>	dinitrogen oxide
10) aluminum cyanide	31) cobalt (III) chromate	10)	
11) Cr <sub>2</sub> O <sub>3</sub>	32) silver cyanide		
12) V <sub>2</sub> O <sub>5</sub>	33) vanadium (V) nitride		
13) Sn <sub>2</sub> O <sub>3</sub>	34) strontium acetate		
14) TiO <sub>2</sub>	35) methylbenzene sulfone		
15) Tl <sub>2</sub> (C <sub>6</sub> H <sub>5</sub> O <sub>2</sub> ) <sub>3</sub>	36) ammonium sulfide		
16) V <sub>2</sub> S <sub>3</sub>			
17) Cr <sub>2</sub> (OH) <sub>3</sub>			
18) Li			
19) Pb <sub>3</sub> N <sub>4</sub>			
20) AgBr			
<b>Ionic/Covalent Compound Naming Solutions</b>		<b>Answers - Naming Chemical Compounds</b>	
1) NaBr	1) Na <sub>2</sub> CO <sub>3</sub>	1) NaF	sodium bromide
2) Sc(OH) <sub>3</sub>	2) O <sub>2</sub>	2) CaCl <sub>2</sub>	calcium acetate
3) V <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	3) NH <sub>3</sub>	3) P <sub>2</sub> O <sub>5</sub>	diphosphorus pentoxide
4) NH <sub>4</sub> F	4) SO <sub>3</sub>	4) Ti(SO <sub>4</sub> ) <sub>3</sub>	titanium(IV) sulfate
5) CdO	5) SO <sub>2</sub>	5) FePO <sub>4</sub>	iron(II) phosphate
6) NiPO <sub>4</sub>	6) CO	6) K <sub>2</sub> N	potassium nitride
7) Li <sub>2</sub> SO <sub>4</sub>	7) O <sub>2</sub>	7) SO <sub>3</sub>	sulfur dioxide
8) Zn <sub>2</sub> P <sub>3</sub>	8) N <sub>2</sub> O <sub>3</sub>	8) Cu <sub>2</sub> H	copper(I) hydroxide
9) Sr(C <sub>6</sub> H <sub>5</sub> O <sub>2</sub> ) <sub>2</sub>	9) CH <sub>4</sub>	9) Zn(NO <sub>3</sub> ) <sub>2</sub>	zinc nitrate
10) Ag <sub>2</sub> O	10) NH <sub>3</sub>	10) V <sub>2</sub> S <sub>3</sub>	vanadium(II) sulfide
11) YCl <sub>3</sub>			
12) YClO <sub>3</sub>			
13) Sn <sub>2</sub> S <sub>3</sub>			
14) Ti <sub>2</sub> N			
15) KMnO <sub>4</sub>			
16) Pb <sub>3</sub> N <sub>4</sub>			
17) cobalt (II) carbonte			
18) CdSO <sub>4</sub>			
19) Cu <sub>2</sub> (NO <sub>3</sub> ) <sub>2</sub>			
20) Fe <sub>2</sub> (CO <sub>3</sub> ) <sub>3</sub>			
21) lithium acetate			

<http://www.chemfiesta.com>

Name each fraction represented by the shaded parts.

- 1)  6) 
- 2)  7) 
- 3)  8) 
- 4)  9) 
- 5)  10) 

CC CW #3 Name: \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_

## 2.3 Carbon Compounds Pages 45-49

### Learn about Carbon

 **Understand** the role of open spaces in molecules.

 **Understand** the relationship between # of bonds of the four groups of neutral elements.

### Learn about Chemistry

**The Chemistry of Carbon** organic chemistry is the study of compounds containing the element carbon. Molecules contain atoms of different elements, which can have different effects. Many of the following elements are in the same group of the periodic table.

 **Understand** how carbon is bonded to form a molecule.

 **Understand** how carbon can bond to other elements.

**Molecules** are a collection of atoms bonded together by covalent or ionic bonds. Molecules are found in many places. It is therefore the study of particles in nature. We will learn what makes them unique.

It is often useful to think of molecules as like a collection of atoms that are bonded together. This is called a molecular model. A molecular model is a representation of a molecule that shows its atoms and bonds.

 **Understand** how molecules are made from atoms and hydrogen atoms. Atoms are the basic building blocks of matter. Hydrogen atoms are the most abundant element in the universe.

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<img alt

le euq s;Am se onatnepoen ed n<sup>3</sup>Äisuf ed otnup le orep ,onatnepoen le euq otla s;Am n<sup>3</sup>Äicillube ed otnup nu eneit onatnep le à N .8 ?onatubosi s;Ärararperp om<sup>3</sup>ÄC;À .7 .racifitsJ .onaporpolcic le euq avitcaer sonem se onatubolcic IE .6 .n<sup>3</sup>Äicaer al arap acimÄuq n<sup>3</sup>Äicauce ebircsE ?onaporp ed n<sup>3</sup>Äicaraperp al arap oirasecen ;Äres odic;À ©Äuq ed oidos ed laS .5 ?Äuq roP;À .dadirucso al ne orolc noc anoiccaer on onatm IE .4 .oiranretauc onobrac nu eneitnec euq ojab s;Am ralucelom osep ed onacla led CAPUI ed erbmon le raD .3 ?onate led aluc@Älom al a onehte le saÄritrevnec om<sup>3</sup>ÄC;À .2 ?oneg<sup>3</sup>Ärdih ed n<sup>3</sup>Äicces al se ©ÄuQ;À .1 .senem;Äxe y sabeurp ne sacram serojem renetbo arap ojabart ed sajoh sal esU .samet sol sodot arap selbinopsid n;Ätse satnugerp noc satiutarg ojabart ed sajoH .sortoson rop sadiregus satnugerp sal noc satnugerp sal rad esavrAS .etneiugis anig;Ap al ne satnugerp sal noc fdp otamrof ne n;Ätse ojabart ed sajoh sal sadoT .setnaidutse sus arap satiutarg selbimirpmi ojabart ed sajoh ragracseD senakLA ed n<sup>3</sup>Äicaraperp al arap acimÄuq ojabart ed ajoH 11 esalC sacram serojem racram a ;Äraduya el y senem;Äxe y sabeurp arap lit<sup>9</sup>À yum ;Äres 11 dradnatS acimÄuq al arap senoiculos y satnugerp noc oyasne etsE .11 odarg le ne senakLA ed n<sup>3</sup>Äicaraperp al arap fdP ne elbimirpmi ojabart ed ajoh etneiugis al a esrirever nebed acimÄuq ed setnaidutse soL 11 esalC fdP teehskroW yrtsimehC 11 esalC senakLA ed n<sup>3</sup>AicaraperP etneicer s;Am nemaxe ed n<sup>3</sup>Ärtap le n<sup>9</sup>Äges 11 radn;Ätse sortseam sol rop sada±Äesid odis nah satnugerp noc erbil olutÄpac led saibas ojabart ed sajoh sal .11 odarg le arap senem;Äxe y sabeurp ne sacram serojem renetbo arap etnemairaid solracitcarp y stnemngissA 11 ssalC yrtsimehC fdP sitarg ragracseD n<sup>3</sup>ÄicaraperP 11 ssalC ESBC ,fdP otamrof ne steehskroW 11 ssalC yrtsimehC ragracseD n à € "Pentane. Click on the link below to download CBSE class 11 preparation of alkanes worksheet A. Objectives after completing this section, you must be able to provide the IUPAC name for any given alkane structure (Kekulà © condensed or short). Draw the Kekulà ©, condensed or short structure of a plane, given its name IUPAC. Key terms ensure that it can define and use in context, the key key below. Study notes The nomenclature IUPAC system aims to ensure that each organic compound has an unoccusted and unequal name. that the name of IUPAC of any compound transmits the structure of that compound to a person familiar with the system. One way to check if the name you have given to an alkane is reasonable is to tell the number of carbon utmos involved by the chosen name. For example, if you named a 3-ethyl-4-methylheptane compound, you have indicated that the compound contains a total of 10 carbon tons à € carbon in an ethyl group, and a carbon time in a methyl group. If you were to check the given structure and find 11 carbon tomes, sabrés that had made a mistake. Maybe the name you should have written was 3â € Ethyl-4,4-dimethylheptane! By appointing Alcanes, a common error of beginner students is a failure to choose the long carbon chain. For example, the correct name for the compound shown below is 3-methylheptane. Remember that each substitute must have a number, and do not forget the prefixes: di, tri, tetra, etc. You should use comas for separate numbers, and to separate numbers and substitutes. Note that 3-methylhexan is a word. Hydrocarbons that do not have functional double or triple vigulous groups are classified as alkanes or cyclokanes, depending on whether the carbon utmos of the molecule are fixed only in chains or also in rings. Although these hydrocarbons do not have functional groups, they constitute the framework in the Functional groups are located in other classes of compounds, and provide an ideal starting point to study and name orgil The alkanes and cyclokanes are also members of a largest class of compounds called Alifatic. In a nutshell, alifatic compounds are compounds that do not incorporate aromatic rings into their molecular structure. The following table lists the names of IUPAC assigned to simple continuous chain alkanes from C-1 to C-10. A "ano" common suffix identifies these compounds as alkanes. The long chain alkanes are well known, and their names can be found in many reference books and text. The methane names through the dean must be memorized, since they constitute the ray of many IUPAC names. Fortunately, common numerical prefixes are used in five or more carbon -tomes names. Table 3.4.1: Simple Alcans without brand Molecular Name Structural Flmula Flores de Molecular Name CH2CH3 CH2H3 75 Some important behavior tendencies and terminología The fesles and structures of these Alcans increase evenly by an increase of CH2. A uniform variation of this type in a series of compounds is called homologous. These phramulas fit into the CNH2N+2 rule. This is also the H/C més highly possible for a stable hydrocarbon. As the H/C proportion in these compounds is maximum, we call them saturated (with hydrogen). Starting with butane (C4H10), and becoming large with more large alkones, we observe the existence of alkaline isía. For example, there are five C6H14 isomers, which are shown below as an abbreviated linen formulas (A a e): although these different compounds have the same molecular fode, only one can be called hexane. How are we going to name others? The IUPAC system first requires names to chains without brand, as has been previously seen, and second that we have names for simple rental groups that can be attached to the chains. Examples of some common rental groups occur in the following table. Keep in mind that the "Aneo" suffix is replaced by "YL" in groups of names. Sämbolo R is used to designate a rented group (not specified). Table 3.4.2: Rent groups Name Group CH3â € "C2H5â € "CH3CH2CH2â € "(CH3) 2CH2â € "CH3CH2CH2CH (CH3) â € "(CH3) 3Câ € "Name Methyl Ethyl Propyl Isopropyl Butyl isobutyl sec-butyl Tert-butyl identify and name groups assigned to this chain. Number of the chain consecutively, beginning at the end close to a substitute group. Designate the location of each substitute group by an appropriate number and name. Place the name, listing groups in literacy order. The prefixes di, tri, tetra etc., used to designate several groups of the same type, are not considered alphabetically. The IUPAC names of the hexane isomers are: a hexane B 2-methylpentano c 3-methylpentane d 2,2-dimethylbutane e 2,3-imimethylbutane The huge halgens are very accommodated, using the names: fluoro (fluoro (fluor F-), Chlorine (Cl-), Bromo (Br-) and iodine (i-). Example 3.4.2: Halomal substitution for example, (CH3) 2ch2ch2ch Br will be called 1-Bromo-3-methylbutane. If the halgen binds to a simple rented group, an alternative name "Halide Remote" can be used. Asä, C2H5 Cl can be called chloroethane (no locator number for a carbon chain) or ethyl chloride is needed. The alkanes can be described by the CNH2N+2 general formula. A rented group is formed by eliminating a hydronale of the alkaline chain and is described by the cnh2n+1 fórmula. The elimination of this hydrogen results in a change of stem of -ane to -yl. Check out the following examples. The same concept is apply to any of the straight straight chain alkaline sel es slitem sol a ,adreiuzqi al edsed anedac al raremun a somaznemoc iS .olitem sopurg sod sol ,riced se ,setneserp sotutitsus sod sol rop sonrapucooper somebed olos ,lanoicnuf opurg n<sup>9</sup>Ägnin eneitnec on olpmeje etse euq a odibeD .elbisop ojab s;Am orem<sup>9</sup>Än le renet ebed etneserp otutitsus reiuqlauc secnote ,selanoicnuf sopurg yah on iS .elbisop ojab s;Am onobrac ed orem<sup>9</sup>Än le renet nebed lanoicnuf opurg nu a sodinu sonobrac soL :2# algeR .sonobrac ohco ed atsnoc y luza ne odatlaser ah es agral s;Am onobrac ed anedac aL .adiutitsus y agral s;Am onobrac ed anedac al rigele rop sonrapucooper somebed olos euq ol rop ,lanoicnuf opurg nu eneitnec euq adiutitsus y agral s;Am onobrac ed anedac al ajilE :1# n<sup>3</sup>Aiculos ed algeR ?otneimiuges ed aluc@Älom al ed erbmon le se l;ÄuC;À .ocit@Äbafla nedro ne soricse n@Ätse selanoicnuf sopurg o/y setneyutitsus sus euq ed eser<sup>9</sup>Ägesa ,etnemiroreta sadad salger saremirp sod sal racipa ed s@Äupsed ,riced sE ;ocit@Äbafla nedro le atneuc ne agneT .elbisop ojab s;Am orem<sup>9</sup>Än le renet ebed etneserp etneyutitsus reiuqlauc secnote ,selanoicnuf sopurg yah on iS .elbisop ojab s;Am onobrac ed orem<sup>9</sup>Än le renet ebed lanoicnuf opurg nu a odinu onobrac nU .opurg .man arap salger serT 3HC81 )2HC( 3HC 24H02C enasociE 3HC71 )2HC( 3HC 04H91C onacedA oN 3HC61 )2HC( 3HC 83H81C onacedatC 0 3HC51 )2HC( 3HC 63H71C enacedatpeH 3HC41 )2HC( 3HC 43H61C 82H31C enacetneP 3HC21 )2HC( 3HC 03H41C enacedarteT 3HC11 )2HC( 3HC 82H31C ANACEDIRT 3HC01 )2HC( 3HC 62H21C ENACEDOD 3HC9 )2HC( 3HC 42H11C enacednU 3HC8 )2HC( 3HC 22H01C onaceD 3HC7 )2HC( 3HC 02H9C enanoN )2HC )2HC( 3HC 61H7C enatpeH 3HC4 )2HC( 3HC 41H6C ONAXEH 3HC3 )2HC( 3HC 21H5C ENATNEP 3HC2 )2HC( 3HC 01H4C onatub 3HC2H3HC 8H3C onaporP 3HC3HC 6H2C onatE 4HC 4H onahtem larutcurse alumr<sup>3</sup>ÄF erbmoN .n<sup>3</sup>Aicaunitnac a albat al ne odanoicroporp ,7 ,y 4 sorem<sup>9</sup>Än sol es orole le y ,onobrac ed n<sup>3</sup>Aicisop adnugues al angisa el es omorb le ,otnat ol roP .orole led setna eneiv omorb le ,etnemacit@ÄbaflA .ocit@Äbafla nedro le atneuc somatiseen ,otcerroc n<sup>3</sup>Aicaremum ed ameuqe le lanoicceles aRaP .anedac al ne atxes y adnugues senoicisop sal napuco selanoicnuf sopurg sol ,ahcered al o adreiuzqi al edsed raremun somigele euq aes ay ,sarbalap sarto nE .n<sup>3</sup>Aicisop atxes al aÄrangisa el es omorb le y n<sup>3</sup>Aicisop adnugues al aÄrangisa es orole le y omorb le ,adreiuzqi al ed sedat raremun iS .alger atse aÄrafstas ahcered al o adreiuzqi al edsed anedac al raremun ,olpmeje etse nE .elbisop ojab s;Am orem<sup>9</sup>Än le renet nebed lanoicnuf opurg nu a sodinu sonobrac soL :2# algeR .sonobrac eteis ed atsnoc y luza ne odatlaser ah es agral s;Am onobrac ed anedac aL .oroIc y omorb ,selanoicnuf opurg sod eneitnec olpmeje etsE .lanoicnuf opurg nu eneitnec euq adiutitsus y agral s;Am onobrac ed anedac al ajilE :1# n<sup>3</sup>Aiculos ed algeR ?otneimiuges ed aluc@Älom al ed erbmon le se l;ÄuC;À onatcolitemid-5,2 se aluc@Älom atse ed erbmon IE .solpmeje setneiugis sol ne oralc s;Am ;Arevlov es otpecnoc etsE .sonobrac sol ed n<sup>3</sup>Aicaremum al a otcepser noc acit@Äbafla aincnedecerp eneit onugrin ,socitn@Ädi nos sotutitsus sod sol euq a odibeD .alger arecret al razilitu ed dadisecen yah on ,olpmeje etse nE :3 algeR .elbisop ajab s;Am n<sup>3</sup>Aicaremum al olitem sopurg sol a ad el otsE .n<sup>3</sup>Aicaunitnac a artseum es omoc onobrac ed anedac al ed ohcered odal le ne azneimoc n<sup>3</sup>Aicaremum al ,alger adnugues al recatsitas arap ,otnat ol roP .5 y 2 sorem<sup>9</sup>Än sol aÄrangisa sel es slitem sol a ,ahcered al ed anedac al raremun a somaznemoc iS le le se @ÄuQ;À onatpehorolc-6-omorb-2 :se aluc@Älom atse ed erbmon IE .onobrac ed n<sup>3</sup>Aicisop atxes al angisa of the follow-up molecule? Solution Rule #1: Choose the longest and most replaced carbon chain containing a functional group. This example contains two functional groups, bromine and chlorine, and a substitute, the methyl group. The longest carbon chain has been highlighted in blue and consists of seven carbons. Rule #2: Carbons attached to a functional group must have the lowest possible carbon number. After taking into account functional groups, any present substitute must have the lowest possible carbon number. This particular example illustrates the principle of a difference point. If we number the chain on the left, the bromine, the metyl group and the chlorine would occupy the second, fifth and sixth positions, respectively. This concept is illustrated in the second drawing below. If we number the chain on the right, the chlorine, the methyl group and the brom would occupy the second, third and sixth positions, respectively, which is illustrated in the first drawing below. The position of the methyl, therefore, becomes a point of difference. In the first drawing, the methyl occupies the third position. To satisfy the second rule, we want to choose the numbering scheme that provides the lowest possible numbering of this substitue. Therefore, the first of the two carbon chains shown below is correct. Therefore, the first numbering scheme is appropriate to use. Once you have determined the correct numbering of carbons, it is often useful to make a list, including functional groups, substitutes and the name of the main chain. Rule #3: After applying the first two rules, take into account the alphabetical order. Alphabetically, the bromine comes before the chlorine. Therefore, the bromine is assigned the second carbon position, and the chlorine is assigned the sixthcarbon. main chain: heptano substitutes: 2-chlor 3-methyl 6-bromo the .onaxehlitemirt-4,3,2 erbmon le aÄrdnet otseupmoc etse ,4 " 3,2 senocisop sal ne olitem ed sotutitsus 3 yah euq odaD .)3-C item remirp le aÄrah euq atseupo n<sup>3</sup>Aiccerid al rev( 2-C olitem remirp le recah arap adreiuzqi a ahcered ed somaremun euq onobrac ed agral s;Am aunitnec anedac ed dutignol anu eneit onacla etsE )b .onaonitem-6-lite-4 se erbmon le y dadiroirp amot )luza n<sup>3</sup>Aicaremum ( adreiuzqi al ne arutcurse al euq ol rop ojab s;Am orem<sup>9</sup>Än le agnet euq ereiug detsu ,ocit@Äbafla nedro le ne otla s;Am se lite le euq odad .4 orem<sup>9</sup>Än le se litem ed otutitsbus le ,adreiuzqi a ahcered ed odnasap )ojor ne orem<sup>9</sup>Än ( ahcered al ed arutcurse al araP .4 orem<sup>9</sup>Än aes olite ed etneutitsus le euq recah arap )luza ne adaremum adreiuzqi al a arutcurse( ahcered a adreiuzqi ed somaremun euq onobrac ed agral s;Am aunitnec anedac ed dutignol anu eneit onacla etsE )a atseupseR .sotseupmoc setneiugis sol ed CAPUI ed sodaucedera serbmon sol raD .onatneplitemid-3,3 se CAPUI ed odaucedera erbmon le euq Äsa 3-C ed litem ed sotutitsus sod eneitT .)3-C ne aÄrttemis al a odibed adreiuzqi a ahcered o ahcered a adreiuzqi ed adaremum res aÄrdop euq( agral s;Am onobrac 5 ed aunitnec anedac ed dutignol anu eneit onacla atsE )c .enatpehlitemid-4,3 ed CAPUI ed odaucedera erbmon nu ad euq ojiferp litemid le y sorem<sup>9</sup>Än sose odnacidni somaÄramall ol euq ol rop ,4 " 3 senocisop sal ne litem ed sotutitsus 2 agnet euq ecah otsE .)ahcered a adreiuzqi ed aremun is 4 res a etnerf 3 elbisop orem<sup>9</sup>Än ronem le renet arap somartnec euq litem ed otutitsbus remirp le renetbo arap adreiuzqi a ahcered ed somaremun euq onobrac ed agral s;Am aunitnec anedac ed dutignol anu eneit onacla etsE )b .enatnep aÄres erbmon us ,5 ed onobrac ed anedac ed dutignol anu noc )selaudividni sonob sol sodot( acram nis anacla anu se arutcurse atse euq odaD )a atseupseR .sotseupmoc setneiugis sol ed CAPUI ed sodaucedera serbmon sol raD enatpehlitem-3-orolc-2-omorb-6 ed soicrejE :se aluc@Älom atse ed This carbon alkaline 6 can be number along different chains (see below) as well as in the opposite directions. This shows the two different chains that can be drawn (do (make First substitute in that chain the minor number). The structure of the left (numbered in blue) is the correct option since it causes more substitutes to be in the continuous chain long (3 vs. 2 in the structure of the right). This would make the name of IUPAC of structure 3-ethyl-2,4-dimethylhexan. (Note how ethyl takes priority over methyl and the di- is not considered for literacy). All the following names represent a compound that has been unduly named. Draw the structure of the name and give the right name of IUPAC for the compounds. a) 1,3-dimethylbutano b) 4-ethyhyntano c) 2-ethyl-3-methylpentano response a) The structure that can be drawn for the incorrect name is shown below on the left. When it is correctly renumed (right structure), the correct name must be 2-methylpentane. b) The structure that can be drawn for the wrong name is shown below on the left. When it is correctly renumed (right structure), notice that the long chain is 6 C and we start the numbering at the end of the right to make Methyl's substitute come out in C-3 (instead of being in C-4 if we numerate the opposite direction) the correct name must be 3-methyl hexane. c) The structure that can be drawn for the wrong name is shown below on the left. When it is correctly renumed (structure to the right), notice that the long chain is 6 c and since this mol © is simí © trica (between the 3 & quot carbon; 4), the numbers can begin from any extreme. In this case, we have 3 & quot carbon methyl substitutes; 4 So the right name is 3,4-dimethyl hexane. All the following names represent a compound that has been appointed improperly. Draw the structure of the name and give the right name of IUPAC for the compounds. a) 2,2-diethylheptane b) 2-propylpentane c) 4,4-diethylbutane response a) the which can be drawn for the wrong name is shown below on the left. When you it properly (structure on the right), notice that the longest chain is now 8 CçÄÄÄs and you have an ethyl substituent at C-3 and a methyl substituent also at C-3 so the proper name is 3- ethyl-3-methyloctane. b) The structure that can be drawn for the improper name is shown below on the left. When you renumber it properly (structure on the right), notice that the longest chain is now 7 CçÄÄÄs and since this molecule is symmetrical (at carbon 4), you can start the numbers from either end. There is a methyl substituent at C-4 so the proper name is 4-methylheptane. c) The structure that can be drawn for the improper name is shown below on the left. When you renumber it properly (structure on the right), going from right to left (to make the ethyl substituent have the lowest number possible), the correct name is 3-ethylhexane . Contributors and Attributions Attributions

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