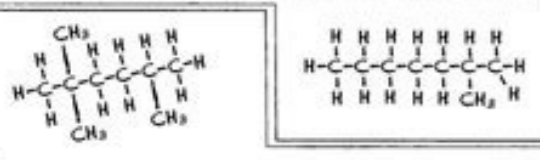


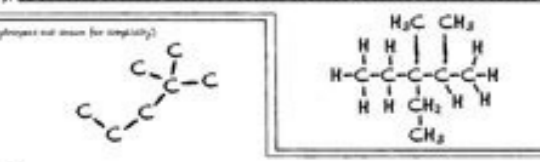
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Name: _____

Systematic (IUPAC) Naming of Alkanes

Directions: Name the following alkanes. Remember, start on the right.

15)  _____

16)  _____

17)  _____

Solutions for the Naming Ionic Compounds Practice Worksheet

- ammonium chloride
- iron (II) nitrate
- barium (II) bromide
- copper (I) phosphide
- tin (IV) selenide
- gallium arsenide
- lead (IV) sulfate
- beryllium bicarbonate
- manganese (II) sulfite
- aluminum cyanide
- Cr_2PO_4
- $\text{V}(\text{O})_2$
- $\text{Sn}(\text{NO}_2)_2$
- Co_2O_3
- $\text{Ti}(\text{C}_2\text{H}_3\text{O}_2)_2$
- V_2S_5
- $\text{Cr}(\text{OH})_3$
- LiI
- Pb_2N_2
- AgBr

- iron (II) phosphate
- titanium (II) selenide
- calcium bromide
- gallium chloride
- sodium hydride
- beryllium hydroxide
- zinc carbonate
- manganese (VII) arsenide
- copper (II) chlorate
- cobalt (III) chromate
- ammonium oxide
- potassium hydroxide
- lead (IV) sulfate
- silver cyanide
- vanadium (V) nitride
- strontium acetate
- molybdenum sulfate
- platinum (II) sulfide
- ammonium sulfate

Ionic/Covalent Compound Naming Solutions

- Na_2CO_3 sodium carbonate
- P_2O_5 diphosphorus pentoxide
- NH_3 ammonia
- FeSO_4 iron (II) sulfate
- SiO_2 silicon dioxide
- GaCl_3 gallium chloride
- CoBr_2 cobalt (II) bromide
- Si_2H_6 silicon tetrahydride
- CO carbon monoxide
- P_4 phosphorus
- dinitrogen trioxide
- nitrogen
- methane
- lithium acetate
- phosphorus trifluoride
- vanadium (V) oxide
- aluminum hydroxide
- zinc sulfide
- silicon tetrafluoride
- silver phosphate

(20) More Naming Practice - Answers

- BBr_3 boron tribromide
- CaSO_4 calcium sulfate
- C_2Br_4 dibromine tetrabromide
- $\text{Cr}(\text{CO})_6$ chromium (VI) carbonyl
- Ag_3P silver phosphide
- IO_2 iodine dioxide
- VO_3 vanadium (V) oxide
- PbS lead (II) sulfide
- CH_4 methane
- N_2O_3 dinitrogen trioxide

Write the formulas of the following chemical compounds:

- tetraphosphorus triselenide P_4Se_3
- potassium acetate $\text{KC}_2\text{H}_3\text{O}_2$
- iron (II) phosphide Fe_3P_2
- diboron hexabromide B_2Br_6
- titanium (IV) nitrate $\text{Ti}(\text{NO}_3)_4$
- dioxmium dioxide Se_2
- copper (I) phosphide Cu_3PO_2
- gallium oxide Ga_2O_3
- tetrastannur dioxide Sn_4O_2
- phosphorus P_4

Answers - Naming Chemical Compounds











- NaBr sodium bromide
- $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ calcium acetate
- P_2O_5 diphosphorus pentoxide
- $\text{Ti}(\text{SO}_4)_2$ titanium(IV) sulfate
- FePO_4 iron(III) phosphate
- K_3N potassium nitride
- SO_2 sulfur dioxide
- Cu_2OH copper(I) hydroxide
- $\text{Zn}(\text{NO}_3)_2$ zinc nitrate
- V_2S_5 vanadium(V) sulfide

Ionic Naming Practice Problems - Solutions

- NaBr sodium bromide
- $\text{Sc}(\text{OH})_3$ scandium hydroxide
- $\text{V}_2(\text{SO}_4)_3$ vanadium (III) sulfate
- NH_4F ammonium fluoride
- CaCO_3 calcium carbonate
- NiPO_4 nickel (II) phosphate
- Li_2SO_4 lithium sulfate
- Zn_3P_2 zinc phosphide
- $\text{Sr}(\text{C}_2\text{H}_3\text{O}_2)_2$ strontium acetate
- Cu_2O copper (I) oxide
- Ag_3PO_4 silver phosphate
- $\text{Y}(\text{ClO}_4)_3$ yttrium chlorate
- SnS_2 tin (IV) sulfide
- $\text{Ti}(\text{CN})_4$ titanium (IV) cyanide
- KMnO_4 potassium permanganate
- Pb_2N_2 lead (II) nitride
- CoCO_3 cobalt (II) carbonate
- CdSO_4 cadmium sulfate
- $\text{Cu}(\text{NO}_2)_2$ copper (II) nitrite
- $\text{Fe}(\text{HCO}_3)_2$ iron (II) bicarbonate
- lithium acetate $\text{LiC}_2\text{H}_3\text{O}_2$

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Name each fraction represented by the shaded parts.

-  _____
-  _____
-  _____
-  _____
-  _____
-  _____
-  _____
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CC CW #3

2.3 Carbon Compounds Pages 45-49

Lesson Objectives

- Describe the physical properties of carbon.
- Describe the structural formulae of each of the first groups of structural isomers.

Lesson Summary

The Chemistry of Carbon (organic chemistry) is the study of compounds with bonds between carbon atoms. Carbon forms four bonds, which allows it to form a wide variety of structures. It is the only element that can form long chains of atoms, and it is the only element that can form rings of atoms. It is the only element that can form both single and double bonds.

- Most carbon compounds are soluble in water and organic solvents.
- Carbon can form structures of different sizes and shapes, such as diamond.

Macromolecules are large molecules that are formed by long chains of atoms that are linked together. They are formed by the reaction of small molecules, such as ethene, to form long chains of atoms. They are formed by the reaction of small molecules, such as ethene, to form long chains of atoms.

Structural Isomers are molecules that have the same molecular formula but different structural formulae. They are formed by the reaction of small molecules, such as ethene, to form long chains of atoms. They are formed by the reaction of small molecules, such as ethene, to form long chains of atoms.

Alkanes are the simplest hydrocarbons, with the general formula $\text{C}_n\text{H}_{2n+2}$. They are formed by the reaction of small molecules, such as ethene, to form long chains of atoms. They are formed by the reaction of small molecules, such as ethene, to form long chains of atoms.

Alkenes are the simplest hydrocarbons with a double bond. They are formed by the reaction of small molecules, such as ethene, to form long chains of atoms. They are formed by the reaction of small molecules, such as ethene, to form long chains of atoms.

Alkynes are the simplest hydrocarbons with a triple bond. They are formed by the reaction of small molecules, such as ethyne, to form long chains of atoms. They are formed by the reaction of small molecules, such as ethyne, to form long chains of atoms.

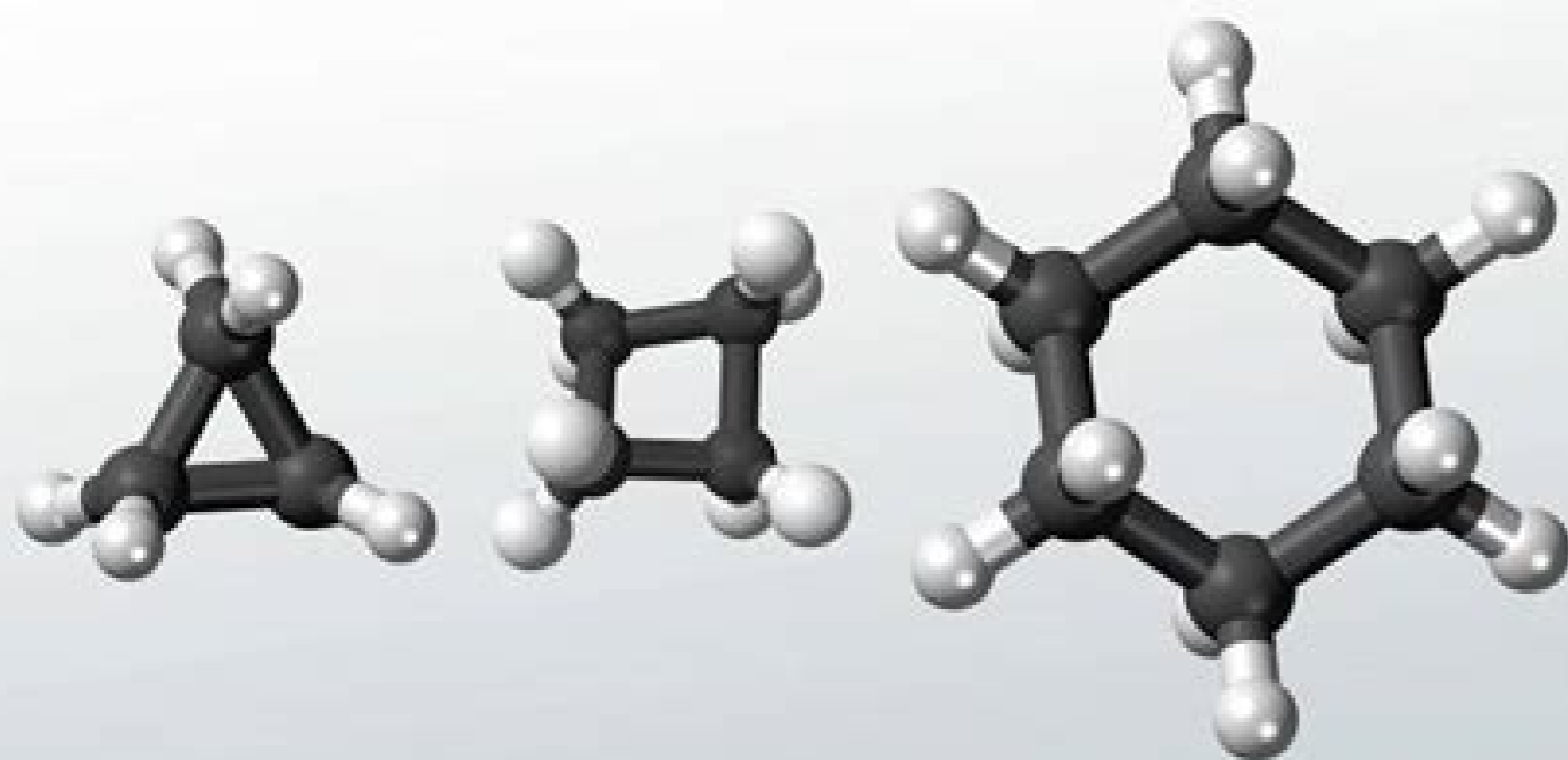
- There are 10 alkanes with the molecular formula $\text{C}_{10}\text{H}_{22}$.
- There are 10 alkenes with the molecular formula $\text{C}_{10}\text{H}_{20}$.
- There are 10 alkynes with the molecular formula $\text{C}_{10}\text{H}_{18}$.

The Chemistry of Carbon

1. How many structural isomers does each of the following have?

a. 10 structural isomers with the molecular formula $\text{C}_{10}\text{H}_{22}$

CYCLOALKANES



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